

For each element, draw the electrons on each orbit for each atom.

<p>H P(+)= N(o)=</p>									<p>He P(+)= N(o)=</p>
<p>Li P(+)= N(o)=</p>	<p>Be P(+)= N(o)=</p>	<p>B P(+)= N(o)=</p>	<p>C P(+)= N(o)=</p>	<p>N P(+)= N(o)=</p>	<p>O P(+)= N(o)=</p>	<p>F P(+)= N(o)=</p>	<p>Ne P(+)= N(o)=</p>		
<p>Na P(+)= N(o)=</p>	<p>Mg P(+)= N(o)=</p>	<p>Al P(+)= N(o)=</p>	<p>Si P(+)= N(o)=</p>	<p>P P(+)= N(o)=</p>	<p>S P(+)= N(o)=</p>	<p>Cl P(+)= N(o)=</p>	<p>Ar P(+)= N(o)=</p>		
<p>K P(+)= N(o)=</p>	<p>Ca P(+)= N(o)=</p>								

Number of protons (+) = ATOMIC NUMBER (top #)

Number of protons (+) & neutrons (o) = ATOMIC MASS (bottom #)

ATOMIC MASS – ATOMIC NUMBER = Number of neutrons (o)

Protons (+) = electrons (-)...2 fit in 1st orbit, 8 in 2nd orbit, 8 in 3rd orbit